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Name _____ Date _____ Period _____

11. Sketch particles in the three states of matter. How close are the particles and how much do they move?

Solid	Liquid	Gas
Really close together	Able to slide past each other.	Really far apart. They do not touch and will fill their container.

12. What is the density of liquid hand soap if 450 mL of the soap has a mass of 130g? Density = $\frac{\text{Mass}}{\text{Volume}}$ = $\frac{130\text{g}}{450\text{mL}}$ = 0.29 g/mL

13. What is the density of a paperclip with a mass of 2.3g and a volume of 1.2 cm³? Density = $\frac{\text{Mass}}{\text{Volume}}$ = $\frac{2.3\text{g}}{1.2\text{cm}^3}$ = 1.92 g/cm³

14. How does density affect whether something will sink or float? The more closely packed the particles in an object are, the more likely it is that they will sink in a solution. However, this also depends upon the density of the solution

15. Define the following:
- Boiling point: from liquid to gas.
 - Melting point: from solid to liquid
 - Freezing point: from liquid or gas to solid
 - Absolute zero: the point at which all particles stop moving. 0 on the Kelvin scale.

16. Classify the following as physical or chemical changes:

- spilling of milk _____
- bending wire _____
- cutting paper _____
- rusting of a nail _____

17. What is viscosity? The resistance of flow from a liquid. Honey is more viscous than water.

18. Identify the following as elements, compounds, or mixtures:

- copper _____ element
- evaporated tea _____ mixture
- sand and water _____ mixture
- calcium carbonate (CaCO₃) _____ compound

19. Atomic Models:

- John Dalton
- List the four postulates of Dalton's Atomic Theory:
 - All atoms are indivisible.

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